

**Константы нестойкости ( $K_{\text{нест}}$ ) некоторых комплексных ионов  
в водных растворах при 298 К**

Схема диссоциации комплексного иона	Константа нестойкости $K_{\text{нест}}$
$[\text{Ag}(\text{NH}_3)_2]^+ \rightleftharpoons \text{Ag}^+ + 2\text{NH}_3$	$9,3 \cdot 10^{-8}$
$[\text{Ag}(\text{NO}_2)_2]^- \rightleftharpoons \text{Ag}^+ + 2\text{NO}_2^-$	$1,8 \cdot 10^{-3}$
$[\text{Ag}(\text{CN})_2]^- \rightleftharpoons \text{Ag}^+ + 2\text{CN}^-$	$1,1 \cdot 10^{-21}$
$[\text{AgBr}_2]^- \rightleftharpoons \text{Ag}^+ + 2\text{Br}^-$	$7,8 \cdot 10^{-8}$
$[\text{AgI}_4]^{3-} \rightleftharpoons \text{Ag}^+ + 4\text{I}^-$	$1,8 \cdot 10^{-14}$
$[\text{AlF}_6]^{3-} \rightleftharpoons \text{Al}^{3+} + 6\text{F}^-$	$1,44 \cdot 10^{-20}$
$[\text{AuCl}_6]^{3-} \rightleftharpoons \text{Au}^{3+} + 6\text{Cl}^-$	$5,0 \cdot 10^{-22}$
$[\text{Au}(\text{CN})_2]^- \rightleftharpoons \text{Au}^{3+} + 2\text{CN}^-$	$5,0 \cdot 10^{-39}$
$[\text{Co}(\text{NH}_3)_6]^{2+} \rightleftharpoons \text{Co}^{2+} + 6\text{NH}_3$	$7,75 \cdot 10^{-6}$
$[\text{Co}(\text{NH}_3)_6]^{3+} \rightleftharpoons \text{Co}^{3+} + 6\text{NH}_3$	$3,0 \cdot 10^{-33}$
$[\text{Co}(\text{CN})_6]^{4-} \rightleftharpoons \text{Co}^{2+} + 6\text{CN}^-$	$1,0 \cdot 10^{-19}$
$[\text{Cu}(\text{NH}_3)_2]^+ \rightleftharpoons \text{Cu}^+ + 2\text{NH}_3$	$2,24 \cdot 10^{-8}$
$[\text{Cu}(\text{NH}_3)_4]^{2+} \rightleftharpoons \text{Cu}^{2+} + 4\text{NH}_3$	$2,14 \cdot 10^{-13}$
$[\text{Cu}(\text{CN})_2]^- \rightleftharpoons \text{Cu}^+ + 2\text{CN}^-$	$1,0 \cdot 10^{-24}$
$[\text{Cu}(\text{CN})_4]^{2-} \rightleftharpoons \text{Cu}^{2+} + 4\text{CN}^-$	$5,0 \cdot 10^{-28}$
$[\text{Fe}(\text{CN})_6]^{3-} \rightleftharpoons \text{Fe}^{3+} + 6\text{CN}^-$	$1,0 \cdot 10^{-31}$
$[\text{Fe}(\text{CN})_6]^{4-} \rightleftharpoons \text{Fe}^{2+} + 6\text{CN}^-$	$1,0 \cdot 10^{-24}$
$[\text{Hg}(\text{CN})_4]^{2-} \rightleftharpoons \text{Hg}^{2+} + 4\text{CN}^-$	$4,0 \cdot 10^{-42}$
$[\text{HgBr}_4]^{2-} \rightleftharpoons \text{Hg}^{2+} + 4\text{Br}^-$	$2,0 \cdot 10^{-22}$
$[\text{HgI}_4]^{2-} \rightleftharpoons \text{Hg}^{2+} + 4\text{I}^-$	$2,48 \cdot 10^{-30}$
$[\text{Ni}(\text{NH}_3)_4]^{2+} \rightleftharpoons \text{Ni}^{2+} + 4\text{NH}_3$	$1,12 \cdot 10^{-8}$
$[\text{Ni}(\text{NH}_3)_6]^{2+} \rightleftharpoons \text{Ni}^{2+} + 6\text{NH}_3$	$1,86 \cdot 10^{-9}$
$[\text{Zn}(\text{NH}_3)_4]^{2+} \rightleftharpoons \text{Zn}^{2+} + 4\text{NH}_3$	$3,46 \cdot 10^{-10}$
$[\text{Zn}(\text{OH})_4]^{2-} \rightleftharpoons \text{Zn}^{2+} + 4\text{OH}^-$	$3,6 \cdot 10^{-14}$